

Mica polytypism: similarities in the crystal structures of coexisting $1M$ and $2M_1$ oxybiotite

TSUTOMU OHTA, HIROSHI TAKEDA AND YOSHIO TAKÉUCHI

Mineralogical Institute, Faculty of Science
University of Tokyo, Hongo, Tokyo 113, Japan

Abstract

The crystal structures of coexisting $1M$ and $2M_1$ oxybiotites were refined by the X-ray diffraction method in order to examine the role of hydrogen in mica structures and polytypism. The unit-layer structures of coexisting $1M$ and $2M_1$ oxybiotites were found to be identical within the limit of the accuracy of the structure refinement. This similarity leads to the conclusion that they are ideally polytypic in spite of the complexity of unit-layer structures such as micas.

The oxybiotites lack hydrogen and instead are enriched in ferric iron. The structural parameters of the tetrahedral and octahedral layers have been found to be remarkably similar to those of the hydrogenated structures. Comparison with hydrogenated biotite suggests that the lack of hydrogen atoms mainly affects the interlayer configuration and consequently causes the interlayer separation to decrease. This feature can be attributed to the fact that interactions between K^+ and H^+ are very extensive in hydrogenated biotites.

Introduction

Mica is one of the layer silicates in which hydrogen plays an important crystal chemical role. The detailed structure analysis of a trioctahedral mica was initiated by Takéuchi and Sadanaga (1959). Since then many reports have been published on the crystal structures of various mica polytypes and of various chemical compositions, and the structural changes due to cation substitutions have been discussed. For example, the dimensional misfit between the octahedral and tetrahedral layers in mica structures was treated by Radoslovich and Norrish (1962). Takéuchi (1965) compared the two brittle micas, margarite and xanthophyllite.

The structural changes of micas due to cation substitutions have been investigated by several workers (Donnay, Donnay and Takeda, 1964; Hazen and Wones, 1972). The crystal structures of micas at high temperature (Takeda and Morosin, 1975) and at high pressure (Takeda, 1977) have also been studied. Although oxygen fugacity is another important factor in crystallization of minerals in nature, its effect especially that of hydrogen on mica structures has not yet been tackled. Takéuchi (1965) suggested that the interaction between OH groups and interlayer cations could possibly be

correlated with the preference of basic polytypes in micas.

Mica structures are based on the stacking of the unit layers which comprise two tetrahedral layers sandwiching an octahedral layer, and an interlayer. Since hydrogen atoms lie nearly directly below or above the interlayer cation in the hole of tetrahedral layer, they should affect the structure of the unit layer when a unit layer stacks upon another layer. It is therefore suggested that the lack of hydrogen atoms in mica crystals also effectively changes the structure. The effect of the presence or absence of hydrogen atoms on a crystal structure is also of crystal chemical interest in general. Structural changes caused by oxidation-reduction reactions coupled with dehydrogenation-hydrogenation reactions may occur but are not well understood.

Micas have many polytypes or polymorphs. Hendricks and Jefferson (1929) studied the stacking mode of layers in micas. Amelinckx and Dekeyser (1953) attributed mica polytypism to a spiral growth mechanism and gave some stacking modes, while Smith and Yoder (1956) derived six basic regular modes of stacking and gave a nomenclature of mica polytypes, $1M$, $2M_1$, $2M_2$, $2O$, $3T$ and $6H$. The stacking sequences of complex mica polytypes were later determined by Ross, Takeda and Wones

Table 3a. (Continued)

- 8 -

H	K	L	FO	FC	H	K	L	FO	FC
6	6	4	9.83	9.94	7	7	3	5.20	5.52
6	6	5	33.63	34.19	7	7	4	13.46	19.23
6	6	6	29.39	30.76	7	7	5	4.95	4.78
6	6	7	8.65	9.52	7	9	0	22.34	23.09
6	6	8	27.56	27.44	7	9	1	15.95	16.77
6	6	9	30.23	29.88	7	9	2	3.48	8.67
6	6	10	32.37	32.46	8	0	0	14.68	15.27
6	6	11	5.37	5.62	8	0	1	24.07	24.76
6	6	12	19.93	19.68	8	0	2	54.94	56.12
6	6	13	13.33	12.47	8	0	3	12.28	12.05
6	6	14	10.75	10.56	8	0	4	5.12	3.96
6	6	15	6.9	9.22	8	0	5	12.92	12.71
6	6	16	8.70	7.61	8	0	0	19.98	20.98
6	6	17	8.73	7.64	8	0	2	13.90	13.73
6	6	18	12.42	12.53	8	0	3	5.58	4.93
6	6	19	13.93	14.63	8	0	4	13.31	13.92
6	6	20	20.76	20.91	8	0	5	10.62	10.64
6	6	21	22.36	22.10	8	0	6	9.78	9.02
6	6	22	6.21	3.98	8	0	7	9.29	10.05
6	6	23	15.89	15.82	8	0	8	8.46	8.79
6	6	24	6.60	6.30	8	0	9	18.14	21.30
6	6	25	19.43	19.81	8	0	10	17.52	19.06
6	6	26	50.43	52.35	8	0	11	37.12	39.09
6	6	27	3.33	9.33	8	0	12		
6	6	28	38.99	38.65	8	0	13		
6	6	29	6.09	5.29	8	0	14		
6	6	30	46.84	48.82	8	0	15		
6	6	31	16.52	16.75	8	0	16		
6	6	32	9.46	8.56	8	0	17		
6	6	33	29.93	29.88	8	0	18		
6	6	34	8.34	8.33	8	0	19		
6	6	35	13.27	12.00	8	0	20		
6	6	36	4.73	6.22	8	0	21		
6	6	37	9.43	9.17	8	0	22		
6	6	38	7.85	8.79	8	0	23		
6	6	39	19.91	20.35	8	0	24		
6	6	40	7.83	6.95	8	0	25		

Table 3b. (Continued)

-12-

H	K	L	FO	FC
8	6	2	37.45	38.43
8	6	4	27.20	28.04
8	6	6	23.68	22.27
9	1	0	30.50	31.21